

**FINAL JEE-MAIN EXAMINATION – JUNE, 2022**

 (Held On Friday 24<sup>th</sup> June, 2022)

TIME : 3 : 00 PM to 6 : 00 PM

**CHEMISTRY**
**TEST PAPER WITH SOLUTION**
**SECTION-A**

1. 120 of an organic compound that contains only carbon and hydrogen gives 330g of CO<sub>2</sub> and 270g of water on complete combustion. The percentage of carbon and hydrogen, respectively are.

(A) 25 and 75                      (B) 40 and 60  
(C) 60 and 40                      (D) 75 and 25

**Official Ans. by NTA (D)**
**Ans. (D)**

- Sol.** Given mass of organic compound = 120

$$\text{mass of CO}_2(\text{g}) = 330 \text{ g}$$

$$\text{mass of H}_2\text{O} (\ell) = 270 \text{ g}$$

$$\text{mass of carbon} = n_{\text{CO}_2} \times 12$$

$$= \frac{330}{44} \times 12 = 90 \text{ g}$$

$$\% \text{ of carbon} = \frac{90}{120} \times 100 = 75\%$$

$$\text{mass of hydrogen} = n_{\text{H}_2\text{O}} \times 2$$

$$= \frac{270}{18} \times 2 = 30 \text{ g}$$

$$\% \text{ of hydrogen} = \frac{30}{120} \times 100 = 25\%$$

2. The energy of one mole of photons of radiation of wavelength 300 nm is

$$\text{(Given : } h = 6.63 \times 10^{-34} \text{ Js, } N_A = 6.02 \times 10^{23} \text{ mol}^{-1}, \text{ } c = 3 \times 10^8 \text{ ms}^{-1}\text{)}$$

(A) 235 kJ mol<sup>-1</sup>                      (B) 325 kJ mol<sup>-1</sup>  
(C) 399 kJ mol<sup>-1</sup>                      (D) 435 kJ mol<sup>-1</sup>

**Official Ans. by NTA (C)**
**Ans. (C)**

- Sol.** Energy of one mole of photons =  $\frac{hc}{\lambda} \times N_A$

$$= \frac{6.63 \times 10^{-34} \times 3 \times 10^8}{300 \times 10^{-9}} \times 6.02 \times 10^{23}$$

$$= 399.13 \times 10^3 \text{ Joule/mole}$$

$$= 399 \text{ kJ / mole}$$

3. The correct order of bond orders of C<sub>2</sub><sup>2-</sup>, N<sub>2</sub><sup>2-</sup> and O<sub>2</sub><sup>2-</sup> is, respectively.

(A) C<sub>2</sub><sup>2-</sup> < N<sub>2</sub><sup>2-</sup> < O<sub>2</sub><sup>2-</sup>                      (B) O<sub>2</sub><sup>2-</sup> < N<sub>2</sub><sup>2-</sup> < C<sub>2</sub><sup>2-</sup>

(C) C<sub>2</sub><sup>2-</sup> < O<sub>2</sub><sup>2-</sup> < N<sub>2</sub><sup>2-</sup>                      (D) N<sub>2</sub><sup>2-</sup> < C<sub>2</sub><sup>2-</sup> < O<sub>2</sub><sup>2-</sup>

**Official Ans. by NTA (B)**
**Ans. (B)**

- Sol.** Species                      Bond order

C<sub>2</sub><sup>2-</sup>                      3

N<sub>2</sub><sup>2-</sup>                      2

O<sub>2</sub><sup>2-</sup>                      1

4. At 25°C and 1 atm pressure, the enthalpies of combustion are as given below:

Substance	H <sub>2</sub>	C(graphite)	C <sub>2</sub> H <sub>6</sub> (g)
$\frac{\Delta_c H^\ominus}{\text{kJmol}^{-1}}$	-286.0	-394.0	-1560.0

The enthalpy of formation of ethane is

(A) +54.0 kJ mol<sup>-1</sup>                      (B) -68.0 kJ mol<sup>-1</sup>  
(C) -86.0 kJ mol<sup>-1</sup>                      (D) +97.0 kJ mol<sup>-1</sup>

**Official Ans. by NTA (C)**
**Ans. (C)**

- Sol.** C<sub>2</sub>H<sub>6</sub>(g) +  $\frac{7}{2}$ O<sub>2</sub>(g) → 2CO<sub>2</sub>(g) + 3H<sub>2</sub>O(ℓ)

$$\Delta_c H(\text{C}_2\text{H}_6) = 2\Delta_f H \text{ CO}_2(\text{g}) + 3\Delta_f H(\text{H}_2\text{O}, \ell)$$

$$- \Delta_f H(\text{C}_2\text{H}_6, \text{g})$$

$$-1560 = 2(-394) + 3(-286) - \Delta_f H(\text{C}_2\text{H}_6, \text{g})$$

$$\Delta_f H(\text{C}_2\text{H}_6, \text{g}) = -86 \text{ kJ/mole}$$

5. For a first order reaction, the time required for completion of 90% reaction is 'x' times the half life of the reaction. The value of 'x' is

$$\text{(Given: } \ln 10 = 2.303 \text{ and } \log 2 = 0.3010)$$

(A) 1.12                      (B) 2.43  
(C) 3.32                      (D) 33.31

**Official Ans. by NTA (C)**
**Ans. (C)**

**Sol.** Given  $t_{0.90} = t_{0.90} = xt_{1/2}$

First order rate constant

$$K = \frac{\ln 2}{t_{1/2}} = \frac{1}{xt_{1/2}} \ln \frac{A_0}{A_0 - A_0 \times \frac{90}{100}}$$

$$\frac{\ln 2}{t_{1/2}} = \frac{\ln 10}{xt_{1/2}}$$

$$x = \frac{\ln 10}{\ln 2} = \frac{2.303}{2.303 \times 0.3010} = 3.32$$

6. Metals generally melt at very high temperature. Amongst the following, the metal with the highest melting point will be

- (A) Hg (B) Ag  
(C) Ga (D) Cs

**Official Ans. by NTA (B)**

**Ans. (B)**

**Sol.** Hg, Ga, Cs are liquid near room temperature But Ag(silver) is solid.

7. Which of the following chemical reactions represents Hall-Heroult Process?

- (A)  $\text{Cr}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Cr}$   
(B)  $2\text{Al}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Al} + 3\text{CO}_2$   
(C)  $\text{FeO} + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$   
(D)  $2[\text{Au}(\text{CN})_2]_{(\text{aq})}^- + \text{Zn}(\text{s}) \rightarrow 2\text{Au}(\text{s}) + [\text{Zn}(\text{CN}_4)]^{2-}$

**Official Ans. by NTA (B)**

**Ans. (B)**

**Sol.** Hall Heroult process is the major industrial process for extraction of aluminium.

8. In the industrial production of which of the following, molecular hydrogen is obtained as a byproduct?

- (A) NaOH (B) NaCl  
(C) Na metal (D)  $\text{Na}_2\text{CO}_3$

**Official Ans. by NTA (A)**

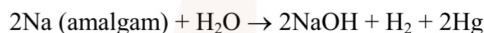
**Ans. (A)**

**Sol.** Sodium hydroxide is generally prepared commercially by electrolysis of sodium chloride in castner Kellner cell.

at cathode :  $\text{Na} + \text{e}^- \xrightarrow{\text{Hg}} \text{Na} - \text{amalgam}$

Anode :  $\text{Cl}^- \longrightarrow \frac{1}{2}\text{Cl}_2 + \text{e}^-$

The Na-amalgam is treated with water to give sodium hydroxide and hydrogen gas :



9. Which one of the following compounds is used as a chemical in certain type of fire extinguishers?

- (A) Baking Soda (B) Soda ash  
(C) Washing Soda (D) Caustic Soda

**Official Ans. by NTA (A)**

**Ans. (A)**

**Sol.** Sodium hydrogencarbonate (Baking soda),  $\text{NaHCO}_3$  is used in the fire extinguishers.

10.  $\text{PCl}_5$  is well known. but  $\text{NCl}_5$  is not. Because.

- (A) nitrogen is less reactive than phosphorous.  
(B) nitrogen doesn't have d-orbitals in its valence shell.  
(C) catenation tendency is weaker in nitrogen than phosphorous.  
(D) size of phosphorous is larger than nitrogen.

**Official Ans. by NTA (B)**

**Ans. (B)**

**Sol.**  $\text{PCl}_5$  forms five bonds by using the d-orbitals to "expand the octet". But  $\text{NCl}_5$  does not exist because there are no d-orbitals in the valence shell ( $2^{\text{nd}}$  shell). Therefore there is no way to expand the octet.

11. Transition metal complex with highest value of crystal field splitting ( $\Delta_0$ ) will be

- (A)  $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$  (B)  $[\text{Mo}(\text{H}_2\text{O})_6]^{3+}$   
 (C)  $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$  (D)  $[\text{Os}(\text{H}_2\text{O})_6]^{3+}$

Official Ans. by NTA (D)

Ans. (D)

Sol. CFSE of octahedral complexes with water is greater for 5d series metal centre ion as compared to 3d and 4d series metal centre.

12. Some gases are responsible for heating of atmosphere (green house effect). Identify from the following the gaseous species which does not cause it.

- (A)  $\text{CH}_4$  (B)  $\text{O}_3$   
 (C)  $\text{H}_2\text{O}$  (D)  $\text{N}_2$

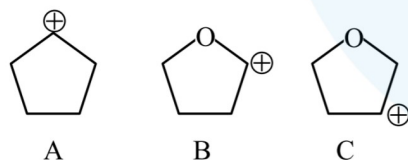
Official Ans. by NTA (D)

Ans. (D)

Sol.  $\text{CH}_4$ ,  $\text{O}_3$  and  $\text{H}_2\text{O}$  causes global warming in Tropospheric level.

$\text{N}_2$  does not cause global warming.

13. Arrange the following carbocations in decreasing order of stability.

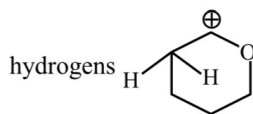


- (A)  $A > C > B$  (B)  $A > B > C$   
 (C)  $C > B > A$  (D)  $C > A > B$

Official Ans. by NTA (A)

Ans. (Bonus)

Sol. Carbocation is stabilised by resonance with lone pairs on oxygen atom and +H effect of 2 $\alpha$



$B > A > C$

14. Given below are two statements.

Statement I : The presence of weaker  $\pi$ - bonds make alkenes less stable than alkanes.

Statement II : The strength of the double bond is greater than that of carbon-carbon single bond.

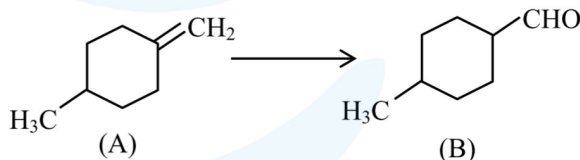
In the light of the above statements, choose the correct answer from the options given below.

- (A) Both Statement I and Statement II are correct.  
 (B) Both Statement I and Statement II are incorrect.  
 (C) Statement I is correct but Statement II is incorrect.  
 (D) Statement I is incorrect but Statement II is correct.

Official Ans. by NTA (A)

Ans. (A)

15. Which of the following reagents/ reactions will convert 'A' to 'B'?

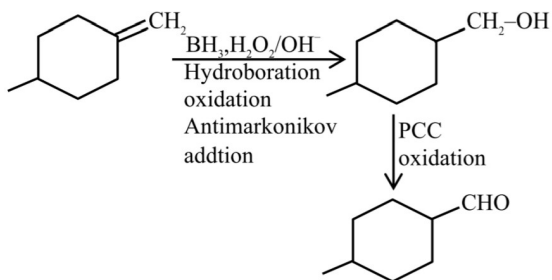


- (A) PCC oxidation  
 (B) Ozonolysis  
 (C)  $\text{BH}_3, \text{H}_2\text{O}_2 / ^-\text{OH}$  followed by PCC oxidation  
 (D)  $\text{HBr}$ , hydrolysis followed by oxidation by  $\text{K}_2\text{Cr}_2\text{O}_7$ .

Official Ans. by NTA (C)

Ans. (C)

Sol.  $\text{BH}_3, \text{H}_2\text{O}_2/\text{OH}$  followed by PCC oxidation.

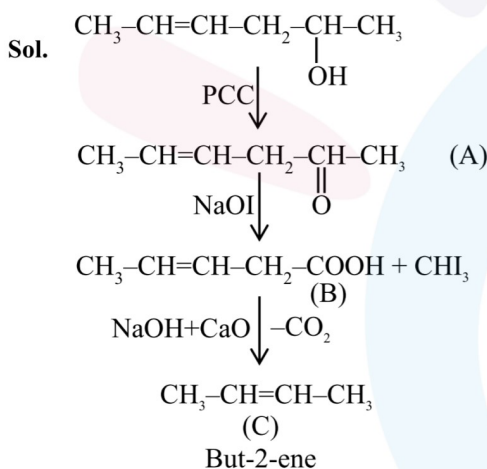


16. Hex-4-ene-2-ol on treatment with PCC gives 'A'. 'A' on reaction with sodium hypoiodite gives 'B', which on further heating with soda lime gives 'C'. The compound 'C' is

- (A) 2-pentene (B) propanaldehyde  
(C) 2-butene (D) 4-methylpent-2-ene

Official Ans. by NTA (C)

Ans. (C)



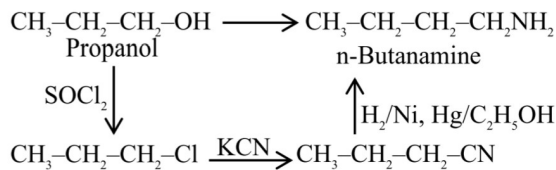
17. The conversion of propan-1-ol to n-butylamine involves the sequential addition of reagents. The correct sequential order of reagents is.

- (A) (i)  $\text{SOCl}_2$  (ii)  $\text{KCN}$  (iii)  $\text{H}_2/\text{Ni, Na(Hg)}/\text{C}_2\text{H}_5\text{OH}$   
(B) (i)  $\text{HCl}$  (ii)  $\text{H}_2/\text{Ni, Na(Hg)}/\text{C}_2\text{H}_5\text{OH}$   
(C) (i)  $\text{SOCl}_2$  (ii)  $\text{KCN}$  (iii)  $\text{CH}_3\text{NH}_2$   
(D) (i)  $\text{HCl}$  (ii)  $\text{CH}_3\text{NH}_2$

Official Ans. by NTA (A)

Ans. (A)

Sol.



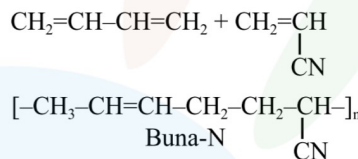
18. Which of the following is **not** an example of a condensation polymer?

- (A) Nylon 6,6 (B) Decron  
(C) Buna-N (D) Silicone

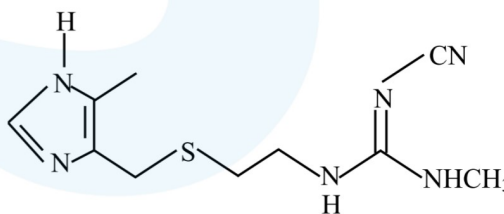
Official Ans. by NTA (C)

Ans. (C)

Sol. Buna-N is an addition copolymer of 1,3-butadiene and acrylonitrile.



19. The structure shown below is of which well-known drug molecule?



- (A) Ranitidine (B) Seldane  
(C) Cimetidine (D) Codeine

Official Ans. by NTA (C)

Ans. (C)

20. In the flame test of a mixture of salts, a green flame with blue centre was observed. Which one of the following cations may be present?

- (A)  $\text{Cu}^{2+}$  (B)  $\text{Sr}^{2+}$   
(C)  $\text{Ba}^{2+}$  (D)  $\text{Ca}^{2+}$

Official Ans. by NTA (A)

Ans. (A)

Sol.	Ion	Colour of the flame
	(A) $\text{Cu}^{+2}$	green flame with blue centre
	(B) $\text{Sr}^{2+}$	Crimson Red
	(C) $\text{Ba}^{2+}$	Apple green

**SECTION-B**

1. At 300 K, a sample of 3.0 g of gas A occupies the same volume as 0.2 g of hydrogen at 200 K at the same pressure. The molar mass of gas A is \_\_\_\_ g  $\text{mol}^{-1}$  (nearest integer) Assume that the behaviour of gases as ideal. (Given: The molar mass of hydrogen ( $\text{H}_2$ ) gas is  $2.0 \text{ g mol}^{-1}$ )

**Official Ans. by NTA (45)**

**Ans. (45)**

**Sol.** Given : Ideal gas A and  $\text{H}_2$  gas at same pressure and volume.

From ideal gas equation  $pV = nRT$

$$n_1 T_1 = n_2 T_2$$

$$\frac{3}{\text{GMM of A}} \times 300 = \frac{0.2}{2} \times 200$$

GMM of A = 45 g/mole

2. A company dissolves 'X' amount of  $\text{CO}_2$  at 298 K in 1 litre of water to prepare soda water

$$X = \_\_\_ \times 10^{-3} \text{ g. (nearest integer)}$$

(Given: partial pressure of  $\text{CO}_2$  at 298 K = 0.835 bar.

Henry's law constant for  $\text{CO}_2$  at 298 K = 1.67 kbar.

Atomic mass of H,C and O is 1, 12 and 6  $\text{g mol}^{-1}$ , respectively)

**Official Ans. by NTA (1221)**

**Ans. (1222 & 1223)**

**Sol.** From Henry law

$$P = K_H X_{\text{CO}_2}$$

$$0.835 = 1.67 \times 10^3 \times 1.67 \times 10^3 \times \frac{w_{\text{CO}_2} / 44}{44 + \frac{1000}{18}}$$

$$w_{\text{CO}_2} = 1.2228 \text{ g} = 1222.8 \times 10^{-3} \text{ g}$$

Or

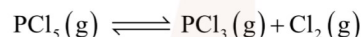
$$P = K_H X_{\text{CO}_2}$$

$$0.835 = 1.67 \times 10^3 \times \frac{n_{\text{CO}_2}}{n_{\text{CO}_2} + n_{\text{H}_2\text{O}}}$$

$$0.835 = 1.67 \times 10^3 \times \frac{w_{\text{CO}_2} / 44}{\frac{1000}{18}}$$

$$w_{\text{CO}_2} = 1.2222 \text{ g} = 1222.2 \times 10^{-3} \text{ g}$$

3.  $\text{PCl}_5$  dissociates as



5 moles of  $\text{PCl}_5$  are placed in a 200 litre vessel which contains 2 moles of  $\text{N}_2$  and is maintained at 600 K. The equilibrium pressure is 2.46 atm. The equilibrium constant  $K_p$  for the dissociation of  $\text{PCl}_5$  is \_\_\_\_  $\times 10^{-3}$ . (nearest integer)

(Given:  $R = 0.082 \text{ L atm K}^{-1} \text{ mol}^{-1}$  : Assume ideal gas behaviour)

**Official Ans. by NTA (1107)**

**Ans. (1107)**

**Sol.** Given : 2 mole of  $\text{N}_2$  gas was present as inert gas.

Equilibrium pressure = 2.46 atm



t = 0	5	0	0
t = Eq <sup>m</sup>	5 - x	x	x

from ideal gas equation

$$PV = nRT$$

$$2.46 \times 200 = (5 - x + x + x + 2) \times 0.082 \times 600$$

$$x = 3$$

$$K_p = \frac{n_{\text{PCl}_3} \times n_{\text{Cl}_2}}{n_{\text{PCl}_5}} \times \left[ \frac{P_{\text{total}}}{n_{\text{total}}} \right]$$

$$\frac{3 \times 3}{2} \times \frac{2.46}{10} = 1.107 = 1107 \times 10^{-3}$$

4. The resistance of conductivity cell containing 0.01 M KCl solution at 298 K is 1750  $\Omega$ . If the conductivity of 0.01 M KCl solution at 298 K is  $0.152 \times 10^{-3} \text{ S cm}^{-1}$ , then the cell constant of the conductivity cell is \_\_\_\_  $\times 10^{-3} \text{ cm}^{-1}$ .

**Official Ans. by NTA (266)**

**Ans. (266)**

**Sol.**  $K = \frac{1}{R} \times \text{cell constant}$

$$0.152 \times 10^{-3} = \frac{1}{1750} \text{ cell constant}$$

$$\text{cell constant} = 266 \times 10^{-3}$$

5. When 200 mL of 0.2 M acetic acid is shaken with 0.6 g of wood charcoal, the final concentration of acetic after adsorption is 0.1 M. The mass of acetic acid adsorbed per gram of carbon is \_\_\_\_\_ g.

**Official Ans. by NTA (2)**

**Ans. (2)**

**Sol.** weight of wood charcoal = 0.6 g

$$\text{Mass of acetic acid adsorbed} = \frac{M_1 V_1 - M_2 V_2}{1000} \times 60$$

$$= \frac{0.2 \times 200 - 0.1 \times 200}{1000} \times 60$$

$$= 1.2 \text{ g}$$

Mass of acetic acid adsorbed per gram of

$$\text{carbon} = \frac{1.2}{0.6} = 2$$

6. (a) Baryte, (b) Galena, (c) Zinc blende and (d) Copper pyrites. How many of these minerals are sulphide based?

**Official Ans. by NTA (3)**

**Ans. (3)**

**Sol.**

(1) Baryte :  $\text{BaSO}_4$

(2) Galena :  $\text{PbS}$

(3) Zinc blende :  $\text{ZnS}$

(4) Copper pyrite :  $\text{CuFeS}_2$

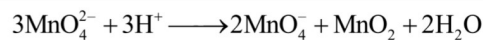
} sulphide ( $\text{S}^{2-}$ ) ores

7. Manganese (VI) has ability to disproportionate in acidic solution. The difference in oxidation states of two ions it forms in acidic solution is \_\_\_\_\_

**Official Ans. by NTA (3)**

**Ans. (3)**

**Sol.**  $\text{MnO}_4^{2-}$  disproportionates in a neutral or acidic solution to give  $\text{MnO}_4^-$  and  $\text{Mn}^{+4}$



O.S. of Mn in  $\text{MnO}_4^- = +7$

O.S. of Mn in  $\text{MnO}_2 = +4$

difference = 3

8. 0.2 g of an organic compound was subjected to estimation of nitrogen by Dumas method in which volume of  $\text{N}_2$  evolved (at STP) was found to be 22.400 mL. The percentage of nitrogen in the compound is \_\_\_\_.[nearest integer]

(Given: Molar mass of  $\text{N}_2$  is  $28 \text{ mol}^{-1}$ . Molar volume of  $\text{N}_2$  at STP : 22.4 L)

**Official Ans. by NTA (14)**

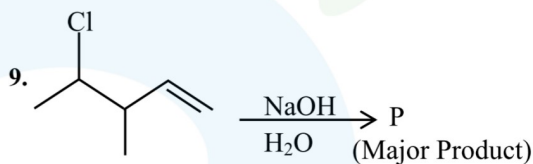
**Ans. (14)**

**Sol.** weight of organic compound = 0.2g

$$\text{mass of } \text{N}_2(\text{g}) \text{ evolved} = \frac{22.4 \times 10^{-3}}{22.4} \times 28$$

$$= 28 \times 10^{-3} \text{ g}$$

$$\% \text{ of N} = \frac{28 \times 10^{-3}}{0.2} \times 100 = 14$$

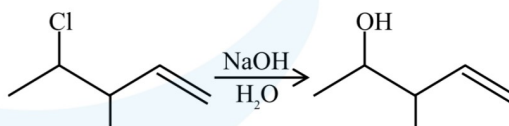


Consider the above reaction. The number of  $\pi$  electrons present in the product 'P' is \_\_\_\_\_.

**Official Ans. by NTA (2)**

**Ans. (2)**

**Sol.** Number of  $\pi$  electron = 2



10. In alanylglycylleucylalanylvaline, the number of peptide linkages is \_\_\_\_\_.

**Official Ans. by NTA (4)**

**Ans. (4)**

**Sol.** There are Five amino acids and four peptide linkages.